Practice Periodic Table: Atomic Number

5

B

11

p=5

n=6

e=5

ve=3

6

C

12

p=

n=

e=

ve=

4

Be

9

p=

n=

e=

ve=

3

Li

7

p=

n=

e=

ve=

Atomic mass\*

2

He

4

p=

n=

e=

ve=

shells=

1

H

1

p=

n=

e=

ve=

shells=

p= 5

n= 6

e= 5

ve= 3

shells= 2

Valence Electrons

u

10

Ne

20

p=

n=

e=

ve=

9

F

19

p=

n=

e=

ve=

8

O

16

p=

n=

e=

ve=

7

N

14

p=

n=

e=

ve=

5

B

11

p=

n=

e=

ve=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

16

S

32

p=

n=

e=

15

P

31

p=

n=

e=

14

Si

28

p=

n=

e=

13

Al

27

p=

n=

e=

17

Cl

35

p=

n=

e=

18

Ar

40

p=

n=

e=

12

Mg

24

p=

n=

e=

11

Na

23

p=

n=

e=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

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shells=

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e=

ve=

shells=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

20

Ca

40

p=

n=

e=

19

K

39

p=

n=

e=

p=

n=

e=

ve=

shells=

p=

n=

e=

ve=

shells=

\*Masses are given for the most common isotopes.